

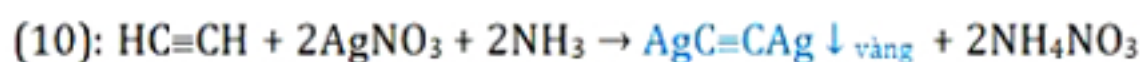
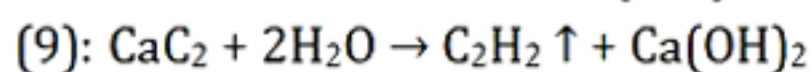
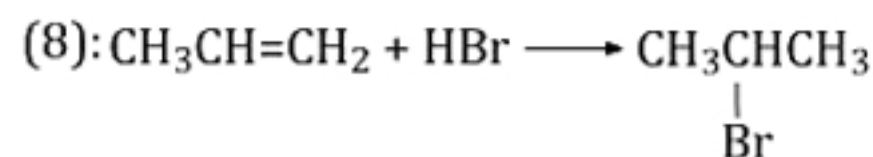
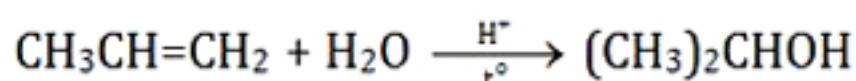
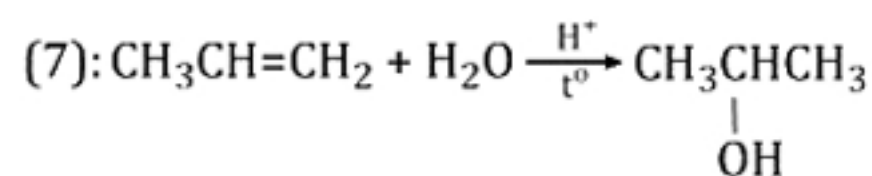
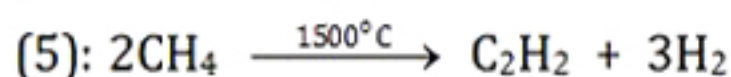
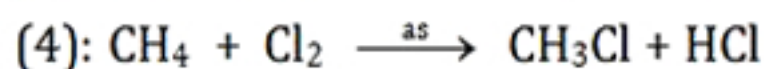
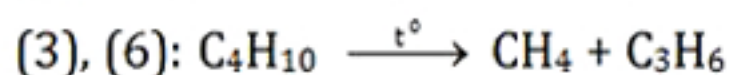
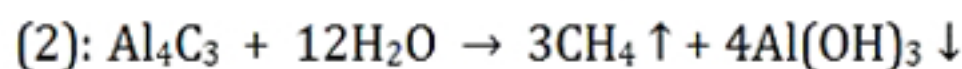
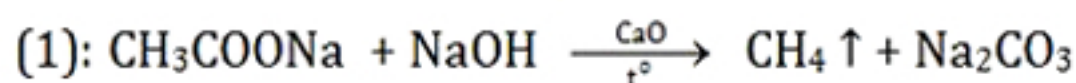


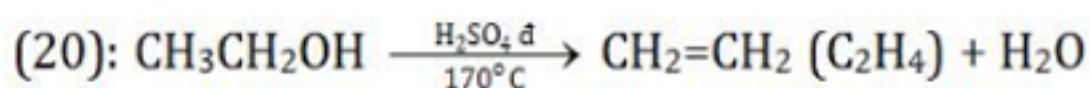
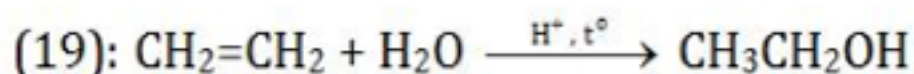
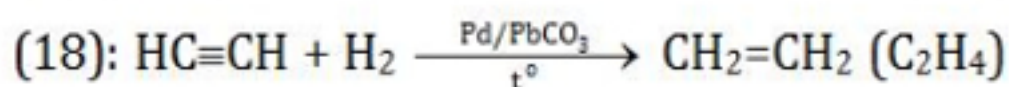
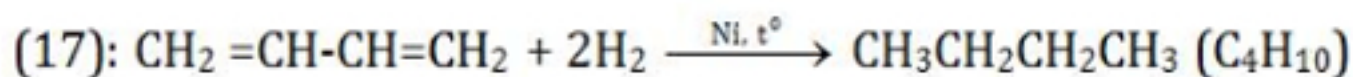
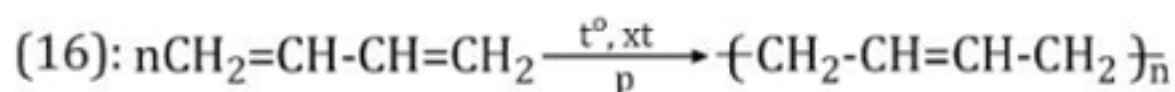
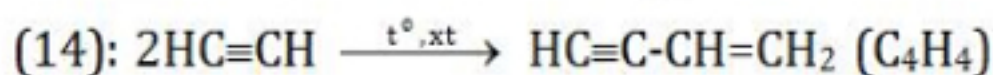
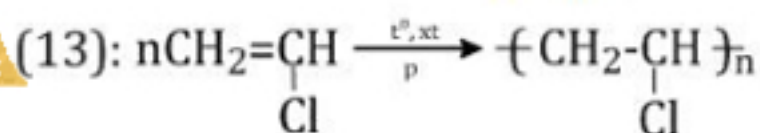
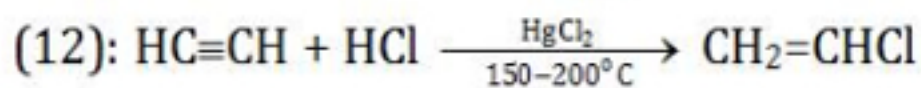
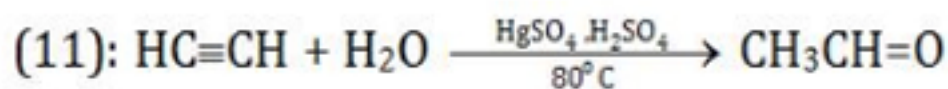
Tổng hợp

Phương Trình Hóa Học

cho team sợ môn hóa
lớp 12

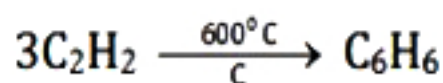
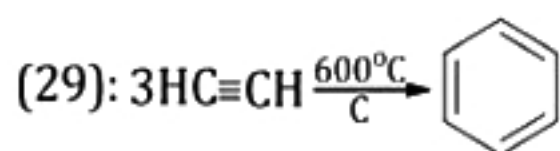
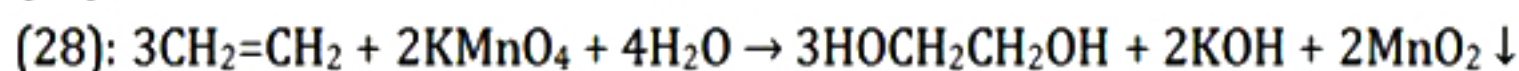
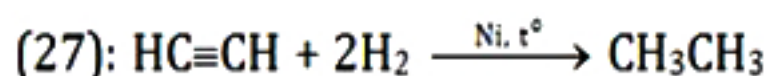
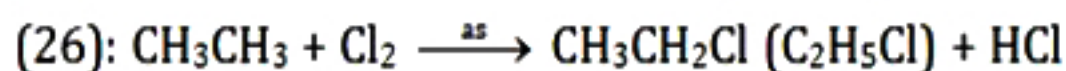
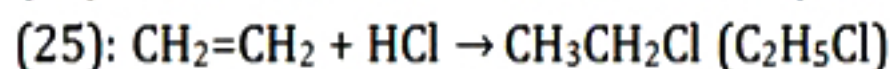
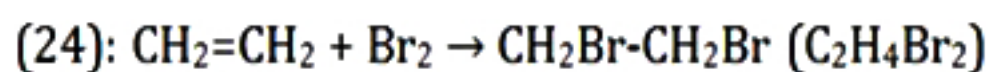
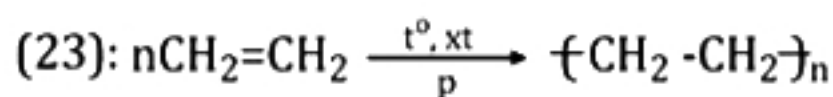
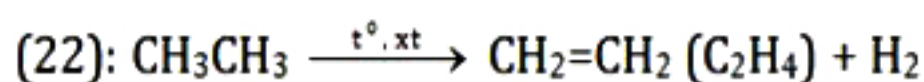
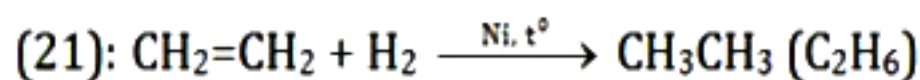


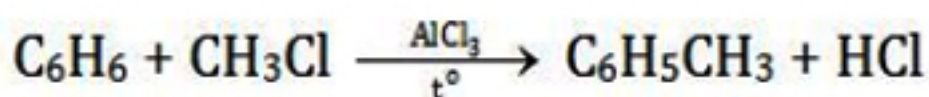
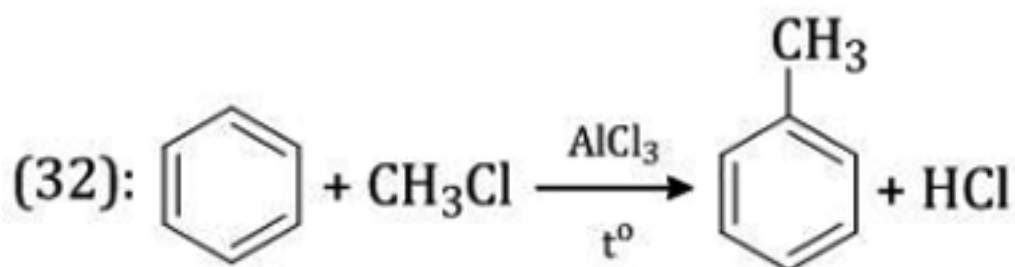
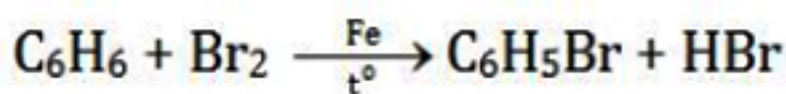
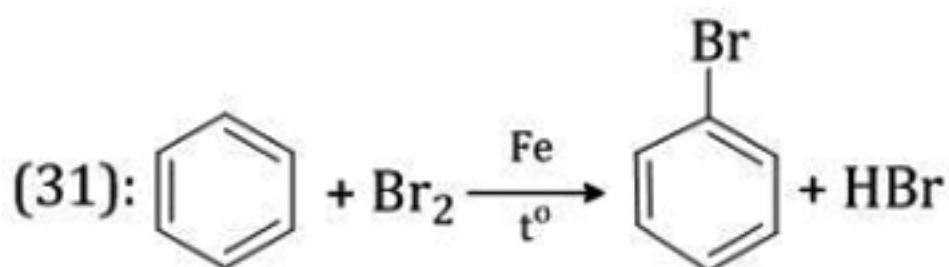
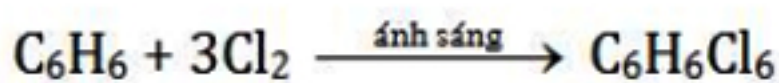
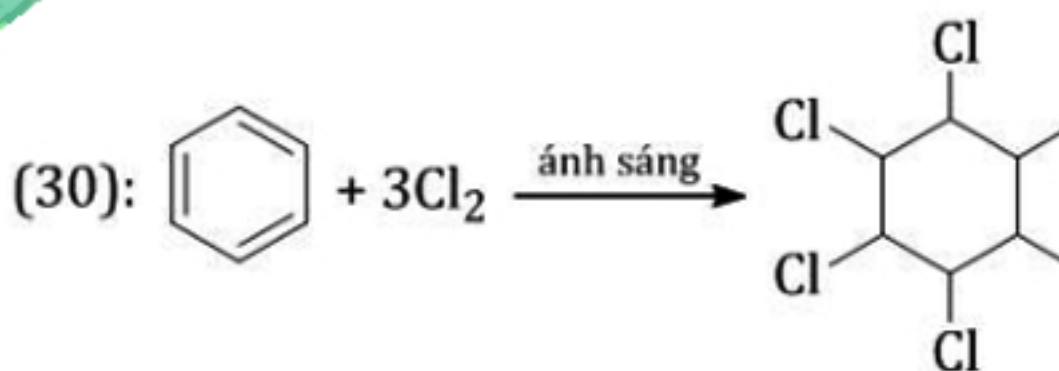


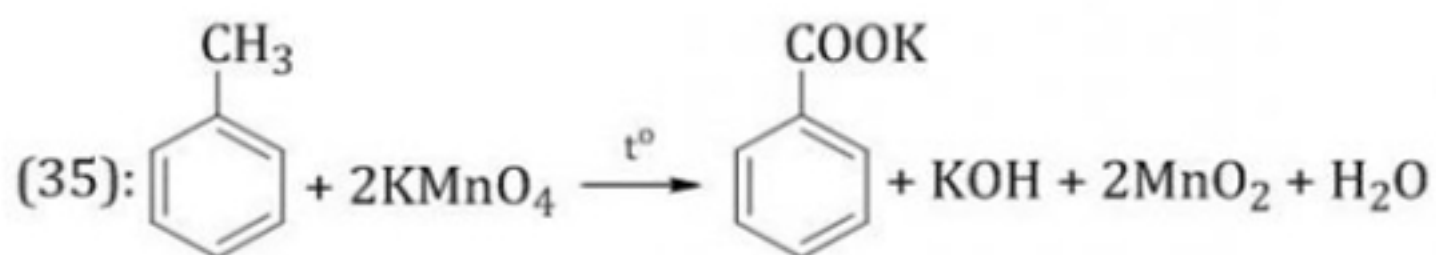
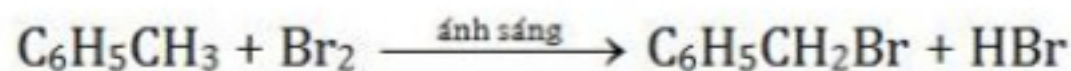
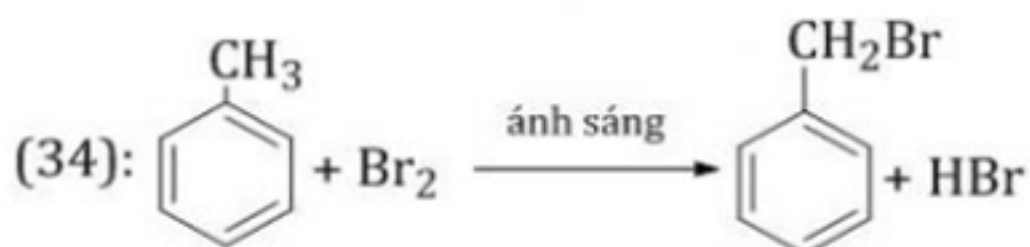
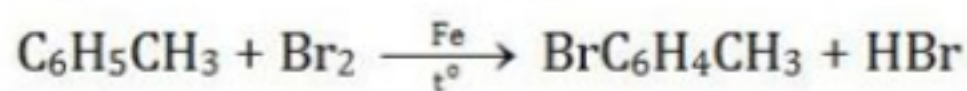
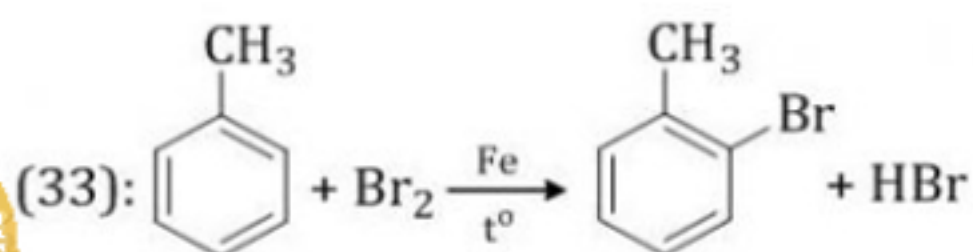


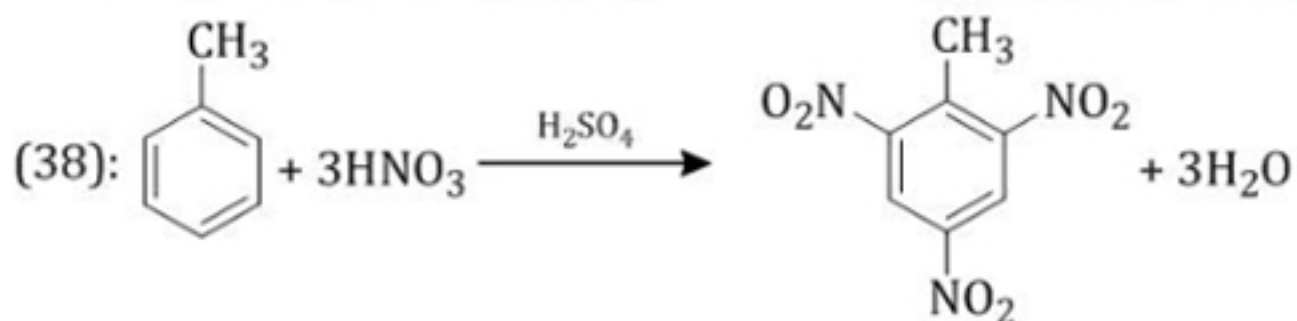
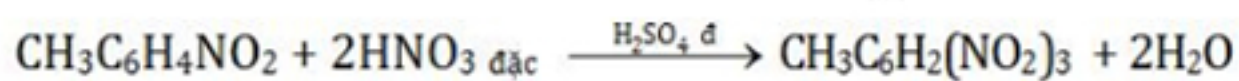
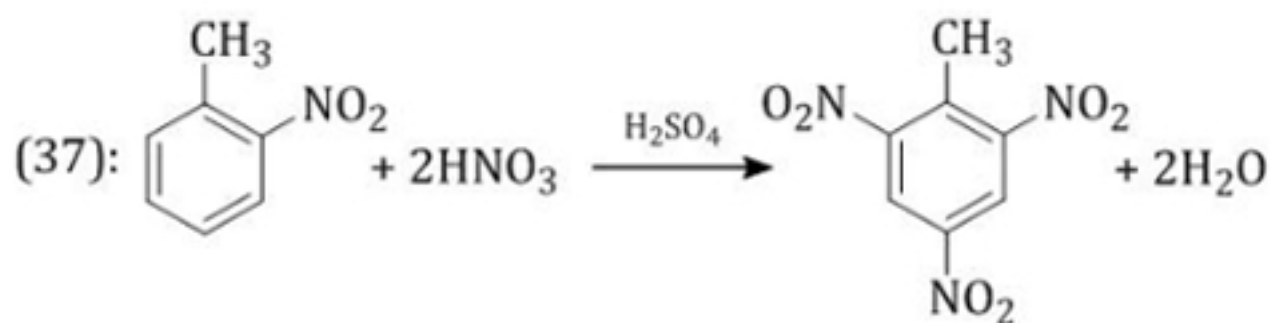
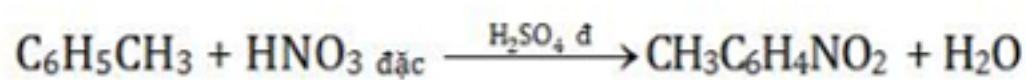
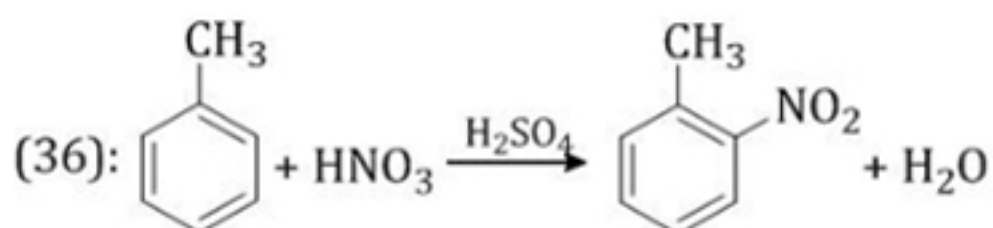
Hi











Phản ứng nhiệt phân của một số muối

1. Muối nitrat

- ✓ Đối với kim loại đứng trước Mg trong dãy hoạt động hóa học:
$$2M(NO_3)_x \xrightarrow{t^o} 2M(NO_2)_x + xO_2$$

(Với những kim loại hóa trị II nhớ đơn giản phân hệ số)
- ✓ Đối với kim loại từ Mg đến Cu trong dãy hoạt động hóa học:
$$4M(NO_3)_x \xrightarrow{t^o} 2M_2O_x + 4xNO_2 + xO_2$$

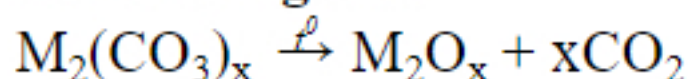
(Với những kim loại hóa trị II nhớ đơn giản phân hệ số)
- ✓ Đối với kim loại đứng sau Cu trong dãy hoạt động hóa học:
$$2M(NO_3)_x \xrightarrow{t^o} 2M + 2xNO_2 + xO_2$$

(Với những kim loại hóa trị II nhớ đơn giản phân hệ số)



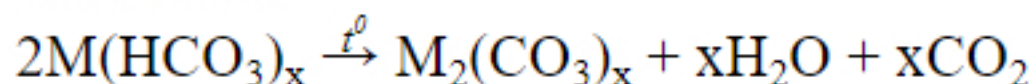
2. Muối cacbonat

✓ **Muối trung hòa:**



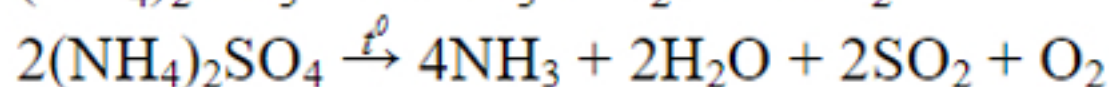
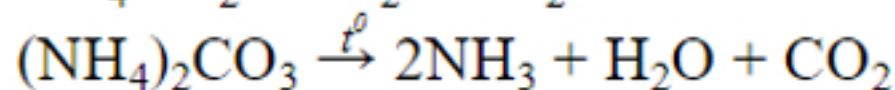
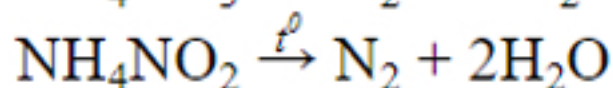
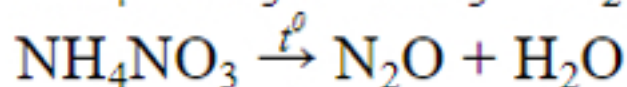
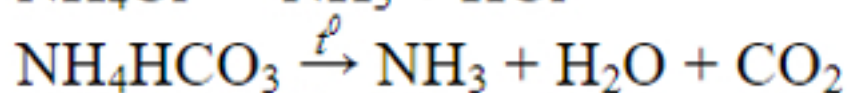
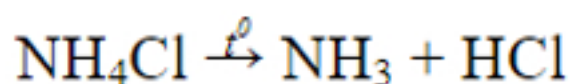
(Với những kim loại hóa trị II nhớ đơn giản phân hệ số)

✓ **Muối axit:**



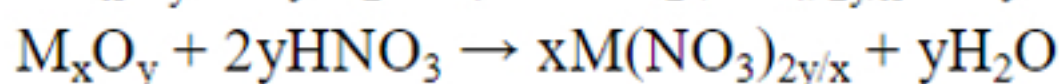
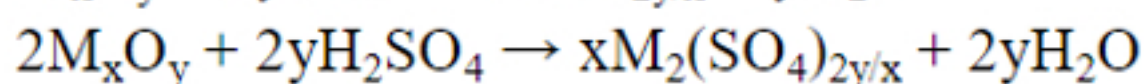
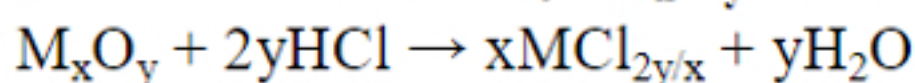
(Với những kim loại hóa trị II nhớ đơn giản phân hệ số)

3. Muối Amoni

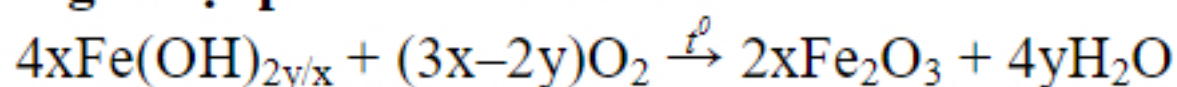


Một số phương trình khi chưa biết hóa trị của kim loại

Hòa tan oxit kim loại M_xO_y vào dung dịch axit:

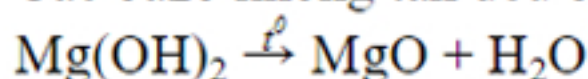


Phản ứng nhiệt phân sắt hiđroxit:



Phản ứng nhiệt phân của các bazơ không tan:

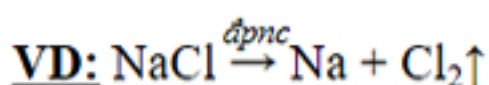
- Các bazơ không tan đều bị nhiệt phân hủy ở nhiệt độ cao:



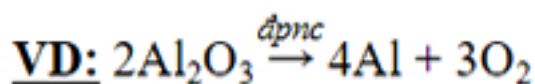


Các phản ứng điều chế một số kim loại

- ✓ Đối với một số kim loại như Na, K, Ca, Mg thì dùng phương pháp điện phân nóng chảy muối clorua.



- ✓ Đối với Al, dùng phương pháp điện phân nóng chảy Al_2O_3 .



- ✓ Đối với các kim loại như Fe, Cu, Pb có thể dùng các phương pháp như:
 - Dùng H_2 : $\text{Fe}_x\text{O}_y + y\text{H}_2 \xrightarrow{t^\circ} x\text{Fe} + y\text{H}_2\text{O}$
 - Dùng C: $2\text{Fe}_x\text{O}_y + y\text{C} \xrightarrow{t^\circ} 2x\text{Fe} + y\text{CO}_2 \uparrow$
 - Dùng CO: $\text{Fe}_x\text{O}_y + y\text{CO} \xrightarrow{t^\circ} x\text{Fe} + y\text{CO}_2 \uparrow$
 - Dùng Al (nhiệt nhôm): $3\text{Fe}_x\text{O}_y + 2y\text{Al} \xrightarrow{t^\circ} 3x\text{Fe} + \text{Al}_2\text{O}_3$